

Chapter 2 The Chemistry Of Life Study Guide Answers

Chapter 1 : Chapter 2 The Chemistry Of Life Study Guide Answers

Chapter 2 the chemistry of life water dissolves polar and ionic acids donate hydrogen nuclei must have a basic appreciation of the principles of chemistry. 2. most of the substances in organisms are based on carbon-containing compounds. the answers to all these questions are based on chemistry. this chapter provides an overview of the Chapter 2: chemistry (!) in many ways, life can be viewed as a complicated chemical reaction. modern models of how life works at all levels typically have at least some aspect of chemistry as a major component or underpinning. therefore.... you must understand chemistry to understand life (and to pass this course)!.Section review 2-2 1. a 2. b 3. c 4. c 5. a 6. b 7. c 8. polarity in a water molecule is caused by an uneven distribution of electrons between the oxygen and hydrogen atoms. 9. the concentration of h ions determines whether a solution is acidic or basic. 10. capillary action is the effect of water rising in a narrow tube against the force of 34 chapter 2 the structure of matter and the chemical elements 2.1 solids, liquids, and gases a chemist's primary interest, as described in chapter 1, is the behavior of matter, but to understand the behavior of matter, we must first understand its internal structure.Chapter 2: chemistry & life 2. molecules 1. atoms 3. water. 1. atoms. a "generic" atom. all matter is made of atoms atoms are the smallest indivisible unit of matter key terms for chapter 236 section summaries/chapter 2 name_____ class_____ date _____ chapter 2 the chemistry of life the basic unit of matter is the atom. the subatomic particles that make up atoms are protons, neutrons, and electrons. protons and neutrons have about the same mass. protons are positively charged particles (+),Chapter 2 – chemistry comes alive i. basic chemistry a. definition of concepts 1. matter: anything that occupies space and has mass a. mass: the amount of material in an object b. weight: varies with gravity 2. energy is the capacity to do work (put matter into motion) a. kinetic energy versus potential energy 1) potential energy: stored energy

Chapter 2: the chemistry of life chapter 2: the chemistry of life 2–2 properties of water mr. m. varco st. joseph high school • like other molecules, water is considered electrically neutral – its 10 protons (+) balance of its 10 electrons (-) • with 8 protons, water's oxygen nucleus attracts electrons moreFigure 2.14a biological molecules are formed from their monomers or units by dehydration synthesis and broken down to the monomers by hydrolysis reactions. + monomers linked by covalent bond (a) dehydration synthesis monomers are joined by removal of oh from one monomer and removal of h from the other at the site of bond formation. monomer 1 Chapter 2 chemistry . the chemical compositions of the body's structures determine their function. i. basic chemistry – chemistry is the scientific discipline concerned with the atomic composition and structure of substances and the reactions they undergo a. matter – anything that occupies space and has mass – composition of all livingChapter key concepts biology resource center biology classzorem 2 chemistry of life 2.1 atoms, ions, and molecules all living things are based on atoms and their interactions. 2.2 properties of water water's unique properties allow life to exist on earth. 2.3 carbon-based molecules carbon-based molecules are the foundation of life. 2.4

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